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From partial molar results (eq. (8.15)), the free energy of the solution is $G_{soln} = nA GA0 + nB GB0 + RT H nA \text{Log}@XA D + nB \text{Log}@XB D + HnA + nB L \text{Log}@PDL GA0 nA + GB0 nB + RT H\text{Log}@XA D nA + \text{Log}@XBD nB + \text{Log}@PD HnA + nBLL$ Because GA0 and GB0 are for constants that do not depend on subsequent increase in temperture, we need to solve for an increase in G_{soln} using $\text{Solve} @ H1 \acute{e} 2L RT H nA \text{Log}@XA D + nB \text{Log}@XB DL == RT H nA \text{Log}@XA D + nB \text{Log}@XB D + HnA + nB L \text{Log}@PDL , PD 99P \text{Ø} E -\text{Log}@XA D nA ...$

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Single-Component Phase Diagrams (cont.); Thermodynamics of Solutions: Gaskell. "The free energy of solution." Chapter 11.5 in Introduction to Metallurgical Thermodynamics. pp. 328-338 E&R: 9.6 and 9.7. L17: X-ray Diffraction: A&T: Study 3.2.2 up to p. 140, and 3.4. Free Energy of Multi-phase Solutions at Equilibrium: E&R: 9.6 and 9.7. L18

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